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AP[®] CHEMISTRY 2010 SCORING GUIDELINES Question 3 (9 points) $8 \text{H}^+(\text{aq}) + 4 \text{Cl}^-(\text{aq}) + \text{MnO}_4^-(\text{aq}) \rightarrow 2 \text{Cl}_2(\text{g}) + \text{Mn}^{2+}(\text{aq}) + 4 \text{H}_2\text{O}(\text{l})$ $\text{Cl}_2(\text{g})$ can be generated in the laboratory by reacting potassium permanganate with an acidified solution of sodium chloride. The net-ionic equation for the reaction is given above. (a) A 25.00 mL sample of 0.250 M

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AP[®] CHEMISTRY 2010 SCORING GUIDELINES (Form B) Question 2 (10 points) $5 \text{Fe}^{2+}(\text{aq}) + \text{MnO}_4^-(\text{aq}) + 8 \text{H}^+(\text{aq}) \rightarrow 5 \text{Fe}^{3+}(\text{aq}) + \text{Mn}^{2+}(\text{aq}) + 4 \text{H}_2\text{O}(\text{l})$ A galvanic cell and the

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balanced equation for the spontaneous cell reaction are shown above. The two reduction half-reactions for the overall reaction that occurs in the cell are shown in the table below.

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AP[®] CHEMISTRY 2010 SCORING GUIDELINES (Form B) Question 1 (10 points) The compound butane, C₄H₁₀, occurs in two isomeric forms, n-butane and isobutane (2-methyl propane). Both compounds exist as gases at 25°C and 1.0 atm. (a) Draw the structural formula of each of the isomers (include all atoms). Clearly label each structure.

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AP[®] CHEMISTRY 2010 SCORING GUIDELINES (Form B) Question 2 (continued) (d) Calculate the value of the equilibrium constant, K_{eq} , for the cell reaction at 25°C. Explain what the magnitude of K_{eq} tells you about the extent of the reaction. $\log K_{eq} = \frac{nE}{5.92}$

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$0.72 = = 61 \cdot 0.0592 \cdot 0.0592$. $K_{eq} = 6.5 \times 10^{60}$ Because the magnitude of K_{eq} is very large, the extent of the cell reaction is also very large and the reaction goes essentially to completion.

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AP® CHEMISTRY 2010 SCORING GUIDELINES Question 1 (10 points) Several reactions are carried out using AgBr, a cream-colored silver salt for which the value of the solubility-product constant, K_{sp} , is 5.0×10^{-13} at 298 K. (a) Write the expression for the solubility-product constant, K_{sp} , of AgBr.

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